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Work is found thethe first law as $w = q - DU$; thus $q = DH$; $w = D HPVL$; 4. Isothermal Process Because U is a function only of T for an ideal gas, $DU = DH = 0$ for an isothermal process. These results also follow from the general results by using $DT = D(PV) = 0$ for an isothermal process.

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The isothermal expansion is conducted at 300 K. $3.4 \Delta H = 42750 \text{ J}$, $\Delta S = 59.7 \text{ J/K}$ 3.5 The final temperature is 323.32 K, which is greater than 323 K because the heat capacity increases with increasing temperature.

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8.54, $c = 0.151$; $C_p AgI = C_p \cdot 8a + 0.3050$, $b = 0$, $c = 0$; $C_p Al = C_p + 20.75 T^2$...

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